

Sta. Lucia High School
Sta. Lucia Dinalupihan Bataan
Third Periodical Test in Science 8

Name:
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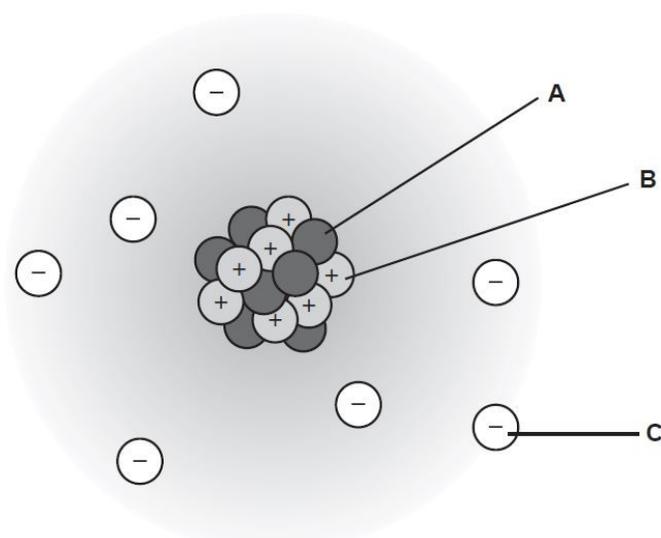
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Multiple Choice. Encircle the letter of the correct answer.

1. Atoms can be viewed under this type of microscope.
 - a. High-powered Light Microscope
 - b. Scanning Tunneling Microscope
 - c. Electron Microscope
 - d. None of these
2. A particle consisting of two or more atoms combined together in a specific arrangement
 - a. Atoms
 - b. particle
 - c. Molecule
 - d. Matter
3. The property of matter which have definite shapes and volume because the properties are packed closely together.
 - a. Solid
 - b. Liquid
 - c. Gas
 - d. Plasma
4. It refers to the amount of material contained in an object.
 - a. Weight
 - b. Mass
 - c. Volume
 - d. none of these
5. Which of the following is NOT matter?
 - a. air
 - b. salt
 - c. light
 - d. smoke
6. These are properties of matter except _____.
 - a. Air
 - b. Solid
 - c. Gas
 - d. Liquid
7. The process by which the molecules on the surface of a liquid change into gas
 - a. Evaporation
 - b. Condensation
 - c. Saturation
 - d. Filtration
8. The process where a gas is changed into a liquid.
 - a. Evaporation
 - b. Condensation
 - c. Saturation
 - d. Filtration
9. This is the organized chart of Elements.
 - a. Periodic Table of Elements
 - b. Periodic Chart of Elements
 - c. Periodic Graph of Elements
 - d. None of these
10. Which of the following best describes an atom?
 - a. protons and electrons grouped together in a random pattern
 - b. protons and electrons grouped together in an alternating pattern
 - c. a core of protons and neutrons surrounded by electrons
 - d. a core of electrons and neutrons surrounded by protons
11. Which of the following is found farthest from the center of an atom?
 - a. Nucleus
 - b. proton
 - c. neutron
 - d. electron
12. Which class of elements best conducts electricity?
 - a. Metals
 - b. Nonmetals
 - c. Metalloids
 - d. Noble (Inert) Gases
13. What happens when a sample of water turns to ice?
 - a. new molecules are formed.
 - b. the mass of the sample is increased.
 - c. the arrangement of the molecules changes.
 - d. energy is absorbed by the molecules.
14. What characteristic of carbon (C) makes it essential to living organisms?
 - a. Carbon forms crystal structures under certain conditions.
 - b. Carbon can exist as a solid, liquid, or gas.
 - c. Carbon bonds in many ways with itself to form chains.
 - d. Carbon exists in radioactive forms.

15. Alkali metals, and halogens are found respectively in Group__ and __ in the Periodic Table.
 a. 1 and 15 b. 2 and 12 c. 1 and 17 d. 1 and 2
16. Atoms of metallic elements tend to
 a. gain electrons and form positive ions
 b. lose electrons and form positive ions
 c. lose electrons and form negative ions
 d. gain electrons and form negative ions
17. Which of the following generally applies to the noble gases?
 a. high ionization energy, low electronegativity, high reactivity
 b. high ionization energy, high electronegativity, high reactivity
 c. low ionization energy, low electronegativity, low reactivity
 d. high ionization energy, low electronegativity, low reactivity
18. Which are found in the nucleus of an atom?
 a. electrons and protons b. neutrons and protons c. neutrons and electrons d. none of these

Use the Diagram below to answer items 19-22



19. The part of an atom labeled **B** is a(an)
 a. proton b. neutron c. electron d. none
20. The part of an atom labeled **C** is a(an)
 a. proton b. neutron c. electron d. none
21. The part of an atom labeled **A** is a(an)
 a. proton b. neutron c. electron d. none
22. Scientists use this model to study atoms
 a. so large b. so small c. easy to see d. so visible

23-30 Categorize the following elements as:

- | |
|--|
| a. Metalloids
b. Nonmetals
c. Halogens
d. Noble Gases |
|--|

- | | |
|-------------|--------------|
| 23. Krypton | 27. Chlorine |
| 24. Bromine | 28. Oxygen |
| 25. Boron | 29. Helium |
| 26. Iodine | 30. Sulfur |

31-35 Refer to the selection below for items 31-35

- a. Dmitri Mendeleev
- b. John Newlands
- c. Johann Dobereiner
- d. Henry Moseley

- 31. English chemist who proposed the Law of Octaves.
- 32. Organized the order of elements in the periodic table according to their atomic numbers.
- 33. He formed the triads of elements with similar properties.
- 34. He discovered that the atomic number closely corresponds to the nuclear charge of the atom.
- 35. He devised the first periodic table.

36. Among the scientists who were responsible for the development of the periodic table are:
I. Dmitri Mendeleev II. Johann Dobereiner and III. John Newlands.
Arrange their names in order of the history of the development of the periodic table.
- a. I, II, III
 - b. II, III, I
 - c. III, I, II
 - d. III, II, I

37. The number of protons in an atom represents the
- a. ionization energy
 - b. atomic number
 - c. electronegativity
 - d. atomic mass

38. Which of the following is an alkali metal?
- a. Ba
 - b. Li
 - c. Pb
 - d. Fe

39. What family of elements includes fluorine and chlorine?
- a. noble gases
 - b. halogens
 - c. alkali metals
 - d. none of these

40. Elements that have some properties of both metals and nonmetals are called
- a. metalloids
 - b. noble gases
 - c. alkali metals
 - d. none of these

41-45 Provide the symbols of the following elements:

- 41. Helium
- 42. Sulfur
- 43. Silicon
- 44. Hydrogen
- 45. Chlorine

46-50. Tell whether the following statement is:

- A. True
- B. False
- C. Maybe
- D. It depends

- _____ 46. Noble gases are generally reactive.
- _____ 47. Alkali metals are very reactive in their elemental form.
- _____ 48. Heat and electricity can easily flow through metals
- _____ 49. It is impossible to predict the properties of any element by its position in the periodic table.
- _____ 50. In an atom, the number of protons is equal to the number of electrons.

Prepared by:

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